Name: _					Lab T	ГА:			
Lab Day	Mon	Tue(am)	Т	ue(pm)	We	ed	Thu(am)	Thur(pm)	Fri
					Gr	ade			
		>	95	>90	>85	>80	>70	<70	
		Report:							

Some Useful Information: $273.15 + {}^{\circ}C = \text{temperature in K}$ 760 mmHg = 1 atm.

Prelaboratory Quiz Score:

Vapor Pressure of Water:

"You are expected to plot a graph of P vT in order to determine the Vapor Pressure of water at the that you performed the experiment. Even in the unlikely event that that temperature is in fact listed!!

T(°C)	P, (mm Hg)						
13.0	11.2	19.0	16.5	25.0	23.8	31.0	33.7
14.0	12.0	20.0	17.5	26.0	25.2	32.0	35.7
15.0	12.8	21.0	18.7	27.0	26.7	33.0	37.7
16.0	13.6	22.0	19.8	28.0	28.3	34.0	39.9
17.0	14.5	23.0	21.1	29.0	30.0	35.0	42.2
18.0	15.5	24.0	22.4	30.0	31.8	36.0	44.6

Data Collection and Calculations:

		Trial 1	Trial 2	Trial 3 (If necessary)
1.	Barometric pressure (in Atmospheres):			
2.	Temperature (in Kelvin):		T	
3.	Mass of aluminum:			
	Moles of Al:			
4.	Initial buret reading:			
	Final buret reading:			
	Volume of H ₂ :			
5.	Moles of H ₂ based on moles of Al:			
	H ₂ Pressure:			
6.	Ideal Gas Constant:			
	Average value:			

Moles of H ₂ based on moles of Al.	6. The Ideal Gas Constant
	•
ost Laboratory Question: Your TA	will not help you with this final question.
	1. 1 1 1
	sodium liquid and nitrogen gas is used in air bags. What volu 26°C is produced when 75.0g of sodium azide (molar mas
nitrogen gas measured at 725 mm of Hg and 2	sodium liquid and nitrogen gas is used in air bags. What volu 26% is produced when 75.0g of sodium azide (molar mas
nitrogen gas measured at 725 mm of Hg and 2	
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Na: 22.99 g.mol⁻¹ N: 14.01 g.mol⁻¹