

	Trial 1	Trial 2	Trial 3 (If necessary)
1. Barometric pressure (in Atmospheres):			
2. Temperature (in Kelvin):			
3. Mass of aluminum:			
Moles of Al:			
4. Initial buret reading:			
Final buret reading:			
Volume of H ₂ :			
5. Moles of H ₂ based on moles of Al:			
H ₂ Pressure:			
6. Ideal Gas Constant:			
Average value:			

For Trial 1: show calculational method for:

5. Moles of H ₂ based on moles of Al.	6. The Ideal Gas Constant
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Post Laboratory Question:

Your TA will not help you with this final question.

The decomposition of sodium azide (NaN₃) solid to sodium liquid and nitrogen gas is used in air bags. What volume of nitrogen gas measured at 725 mm of Hg and 26 °C is produced when 75.0g of sodium azide (molar mass = 65.01g/mol) decomposes.

Na: 22.99 g.mol⁻¹

N: 14.01 g.mol⁻¹