

SID

Last _____

First _____

Question 1

7 Points

From the following vapor pressure data for iodomethane, calculate the molar heat of vaporization of CH_3I

$$P = 100 \text{ mm Hg @ } 266\text{K}$$

$$P = 400 \text{ mm Hg @ } 298\text{K}$$

Must Show Work for Full Credit - $R = 8.314 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$

 kJ·mol⁻¹**Question 2**

3 Points

What type(s) of intermolecular forces are expected between CH_3NH_2 molecules?

Circle all those that apply.

- | | |
|--|---|
| <input type="checkbox"/> Ion - Ion | <input type="checkbox"/> Ion - Dipole |
| <input type="checkbox"/> Dipole - Dipole | <input type="checkbox"/> Hydrogen bonding |
| <input type="checkbox"/> Induced Dipole - Induced Dipole | |