

SID

Last _____

First _____

Question 1

5 Points

An aqueous solution is 16.5% by mass silver(I) nitrate. What is the mole fraction of silver(I) nitrate in the solution?

Must Show Work for Full Credit**Use the following Atomic Weights (g.mol⁻¹):- Ag 107.87; N 14.01; O 16.00; H 1.01****Question 2**

3 Points

Match the following aqueous solutions with the appropriate letter from the column on the right. Assume complete dissociation of electrolytes.

 0.25 m CrSO₄ 0.15 m CuCl₂ 0.12 m Fe(NO₃)₃ 0.44 m Glucose (nonelectrolyte)

A. Highest boiling point

B. Second highest boiling point

C. Third highest boiling point

D. Lowest boiling point

Question 3

2 Points

The Vapor Pressure of 4 substances was measured at 25°C and they were found to be 143.0 mmHg, 67.9 mm Hg, 151.7 mmHg, 514.4 mmHg. The four substances measured are given below. Which one of the four would you anticipate having the Vapor Pressure of 151.7 mm Hg?

 CH₃OH C₅H₁₂ C₆H₁₄ CH₃CH₂OH