Question 1 5 Points  An aqueous solution is 16.5% by mass silver(T) nitrate. What is the mole fraction of silver(T) nitrate in the solution?  Must Show Work for Full Credit Use the following Atomic Weights (g.mol <sup>-1</sup> ):- Ag 107.87; N 14.01; O 16.00; H 1.01  West the following Atomic Weights (g.mol <sup>-1</sup> ):- Ag 107.87; N 14.01; O 16.00; H 1.01  An aqueous solutions with the appropriate letter from the column on the right. Assume complete dissociation of electrolytes.  O.25 m CrSO4  O.15 m CuCl2  B. Second highest boiling point  O.12 m Fe(NO <sub>3</sub> ) <sub>3</sub> C. Third highest boiling point  D. Lowest boiling point  Question 3  Points  The Vapor Pressure of 4 substances was measured at 25°C and they were found to be 143.0 mmHg, 67.9 mm Hg, 151.7 mmHg, 514.4 mmHg  The four substances measured are given below. Which one of the four would you anticipate having the Vapor Pressure of 151.7 mm Hg?  CH <sub>3</sub> CH <sub>4</sub> CH <sub>3</sub> CH <sub>4</sub> CH <sub>3</sub> CH <sub>2</sub> CH <sub>3</sub> CH <sub>2</sub> CH	Chem 112		Spring 2020	Qui	z 2		Whelan
Silver(I) nitrate in the solution?  Must Show Work for Full Credit  Use the following Atomic Weights (g.mol <sup>-1</sup> ):- Ag 107.87; N 14.01; O 16.00; H 1.01  The right. Assume complete dissociation of electrolytes.  O.25 m CrSO4  O.15 m CrSO4  O.15 m CuCl2  O.15 m CuCl2  O.12 m Fe(NO <sub>3</sub> ) <sub>3</sub> C. Third highest boiling point  D. Lowest boiling point  O.44 m Glucose (nonelectrolyte)  D. Lowest boiling point  Cuestion 3  Points  The Vapor Pressure of 4 substances was measured at 25°C and they were found to be 143.0 mmHg, 67.9 mm Hg, 151.7 mmHg, 514.4 mmHg  The four substances measured are given below. Which one of the four would you anticipate having the Vapor Pressure of 151.7 mm Hg?  CH <sub>3</sub> OH  Ch <sub>3</sub> OH	SID		Last		Firs	+	
Question 2  3 Points  Outs the following aqueous solutions with the appropriate letter from the column on the right. Assume complete dissociation of electrolytes.  Outs m CrSO4  Outs m CrSO4  Outs m Fe(NO3)3  O	•	· ·	•	s silver(I) nitr	ate. W		
the right. Assume complete dissociation of electrolytes.  O.25 m CrSO4  O.15 m CuCl2  B. Second highest boiling point  O.12 m Fe(NO3)3  C. Third highest boiling point  D. Lowest boiling point  D. Lowest boiling point  The Vapor Pressure of 4 substances was measured at 25°C and they were found to be 143.0 mmHg, 67.9 mm Hg, 151.7 mmHg, 514.4 mmHg  The four substances measured are given below. Which one of the four would you anticipate having the Vapor Pressure of 151.7 mm Hg?  CH3OH  CH3OH			se the following Atomic W	/eights (g.mol <sup>-1</sup> ):	- Ag 10		
O.15 m CuCl <sub>2</sub> B. Second highest boiling point  O.12 m Fe(NO <sub>3</sub> ) <sub>3</sub> C. Third highest boiling point  D. Lowest boiling point  D. Lowest boiling point  Cuestion 3  Points  The Vapor Pressure of 4 substances was measured at 25°C and they were found to be 143.0 mmHg, 67.9 mm Hg, 151.7 mmHg, 514.4 mmHg  The four substances measured are given below. Which one of the four would you anticipate having the Vapor Pressure of 151.7 mm Hg?  CH <sub>3</sub> OH	•		•	• •	•	letter from the	column on
0.12 m Fe(NO <sub>3</sub> ) <sub>3</sub>		0.25 m	CrSO4	<b>A</b> .	Highes	t boiling point	нин
O.44 m Glucose (nonelectrolyte)  D. Lowest boiling point  The Vapor Pressure of 4 substances was measured at 25°C and they were found to be 143.0 mmHg, 67.9 mm Hg, 151.7 mmHg, 514.4 mmHg  The four substances measured are given below. Which one of the four would you anticipate having the Vapor Pressure of 151.7 mm Hg?  CH <sub>3</sub> OH  CG <sub>6</sub> H <sub>14</sub>		0.15 m	CuCl <sub>2</sub>	В.	Second	<b>l highest</b> boiling	point
Question 3  2 Points  The Vapor Pressure of 4 substances was measured at $25^{\circ}C$ and they were found to be 143.0 mmHg, 67.9 mm Hg, 151.7 mmHg, 514.4 mmHg  The four substances measured are given below. Which one of the four would you anticipate having the Vapor Pressure of 151.7 mm Hg? $CH_3OH$ $C_6H_{14}$		0.12 m	Fe(NO3)3	<b>C</b> .	Third	<b>highest</b> boiling p	oint
2 Points  143.0 mmHg, 67.9 mm Hg, 151.7 mmHg, 514.4 mmHg  The four substances measured are given below. Which one of the four would you  anticipate having the Vapor Pressure of 151.7 mm Hg?  □ CH <sub>3</sub> OH  □ C <sub>6</sub> H <sub>14</sub>		0.44 m	Glucose (nonelectroly	te) D.	Lowest	boiling point	ппининини
	•	143.0 mmHg, The four substa	<b>67.9 mm Hg</b> , nces measured are giv	151.7 mm en below. Wh	n <b>Hg</b> , ich one	514.	4 mmHg
□ C <sub>5</sub> H <sub>12</sub> □ CH <sub>3</sub> CH <sub>2</sub> OH		☐ CH <sub>3</sub> OH			C <sub>6</sub> H <sub>14</sub>		
		□ <b>C</b> <sub>5</sub> H <sub>12</sub>			CH₃CH	₂OH	111111111111111111111111111111111111111