

Ions: Basic Nomenclature

1A															7A	8
H ⁺	2A														H ⁻	
Li ⁺															F ⁻	
Na ⁺	Mg ²⁺														Cl ⁻	
K ⁺	Ca ²⁺	3B	4B	5B	6B	7B	8B	8B	8B	1B	2B				Br ⁻	
Rb ⁺	Sr ²⁺														I ⁻	
Cs ⁺	Ba ²⁺															

Monatomic Anions

H ⁻	Hydride	O ²⁻	Oxide	N ³⁻	Nitride
F ⁻	Fluoride	S ²⁻	Sulfide	P ³⁻	Phosphide
Cl ⁻	Chloride	Se ²⁻	Selenide		
Br ⁻	Bromide	Te ²⁻	Telluride		
I ⁻	Iodide				

Monatomic Cations

Name is the same as the neutral atom. Memorize the charges on of all bar the 'XB' ones. These are the transition metals and they can have varying charges associated with them.

Polyatomic Anions

CN ⁻	Cyanide	NO ₂ ⁻	Nitrite	OH ⁻	Hydroxide
CH ₃ CO ₂ ⁻	Acetate	NO ₃ ⁻	Nitrate	SO ₃ ²⁻	Sulfite
CO ₃ ²⁻	Carbonate	PO ₄ ³⁻	Phosphate	SO ₄ ²⁻	Sulfate
HCO ₃ ⁻	Hydrogen carbonate	HPO ₄ ²⁻	Hydrogen phosphate	HSO ₄ ⁻	Hydrogen sulfate
		H ₂ PO ₄ ⁻	Dihydrogen phosphate		
ClO ⁻	Hypochlorite	CrO ₄ ²⁻	Chromate		
ClO ₂ ⁻	Chlorite	Cr ₂ O ₇ ²⁻	Dichromate		
ClO ₃ ⁻	Chlorate	MnO ₄ ⁻	Permanganate		
ClO ₄ ⁻	Perchlorate				

Polyatomic Cations

NH₄⁺ Ammonium

Strong Acids

HCl	Hydrochloric acid	HI	Hydroiodic acid	HClO ₄	Perchloric acid
HBr	Hydrobromic acid	HNO ₃	Nitric Acid	H ₂ SO ₄	Sulfuric acid

Strong Bases

LiOH	Lithium hydroxide	NaOH	Sodium hydroxide	KOH	Potassium hydroxide
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