

Ions: Basic Nomenclature

1A																			7A	8
H ⁺	2A																		H ⁻	
Li ⁺																			F ⁻	
Na ⁺	Mg ²⁺	3B	4B	5B	6B	7B	8B	8B	8B	1B	2B	Al ³⁺		N ³⁻	O ²⁻			Cl ⁻		
K ⁺	Ca ²⁺														Se ²⁻			Br ⁻		
Rb ⁺	Sr ²⁺														Te ²⁻			I ⁻		
Cs ⁺	Ba ²⁺																			

Monatomic Anions

H ⁻	Hydride	O ²⁻	Oxide	N ³⁻	Nitride
F ⁻	Fluoride	S ²⁻	Sulfide	P ³⁻	Phosphide
Cl ⁻	Chloride	Se ²⁻	Selenide		
Br ⁻	Bromide	Te ²⁻	Telluride		
I ⁻	Iodide				

Monatomic Cations

Name is the same as the neutral atom.

Polyatomic Anions

CN ⁻	Cyanide	NO ₂ ⁻	Nitrite	OH ⁻	Hydroxide
CH ₃ CO ₂ ⁻	Acetate	NO ₃ ⁻	Nitrate	SO ₃ ²⁻	Sulfite
CO ₃ ²⁻	Carbonate	PO ₄ ³⁻	Phosphate	SO ₄ ²⁻	Sulfate
HCO ₃ ⁻	Hydrogen carbonate	HPO ₄ ²⁻	Hydrogen phosphate	HSO ₄ ⁻	Hydrogen sulfate
		H ₂ PO ₄ ⁻	Dihydrogen phosphate		
ClO ⁻	Hypochlorite	CrO ₄ ²⁻	Chromate		
ClO ₂ ⁻	Chlorite	Cr ₂ O ₇ ²⁻	Dichromate		
ClO ₃ ⁻	Chlorate	MnO ₄ ⁻	Permanganate		
ClO ₄ ⁻	Perchlorate				

Polyatomic Cations

NH₄⁺ Ammonium

Strong Acids

HCl	Hydrochloric acid	HI	Hydroiodic acid	HClO ₄	Perchloric acid
HBr	Hydrobromic acid	HNO ₃	Nitric Acid	H ₂ SO ₄	Sulfuric acid

Strong Bases

LiOH	Lithium hydroxide	NaOH	Sodium hydroxide	KOH	Potassium hydroxide
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