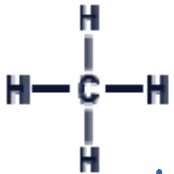
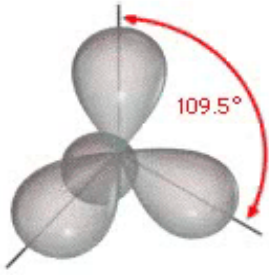
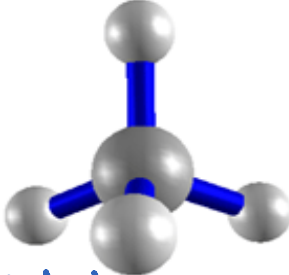

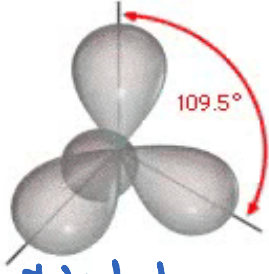
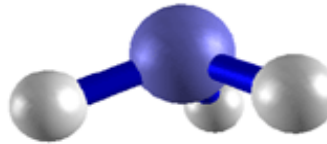

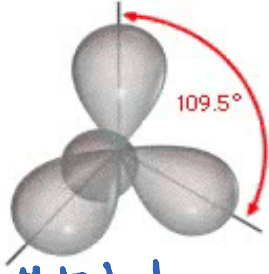
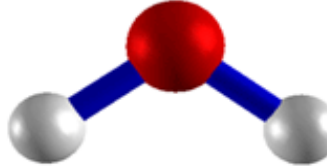

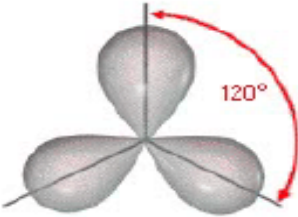
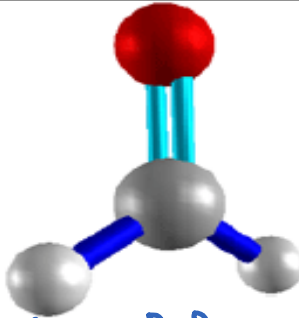
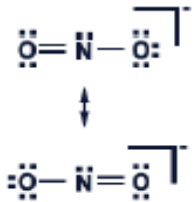
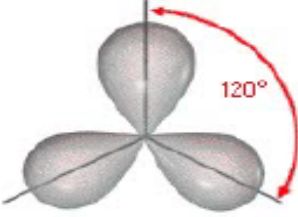
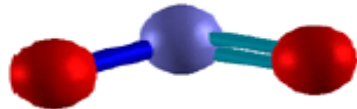
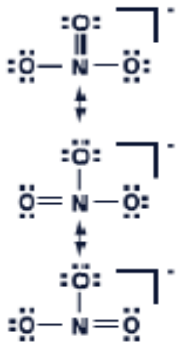
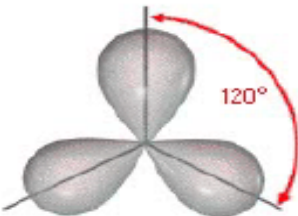
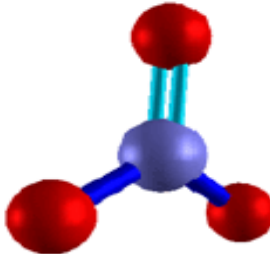


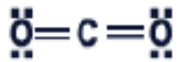
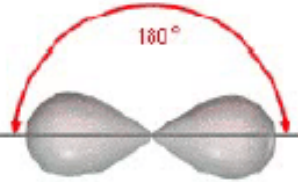

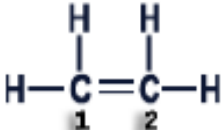
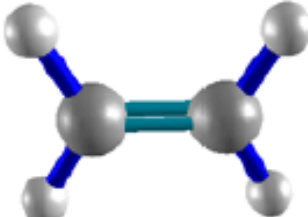
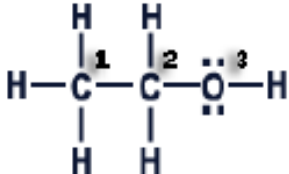
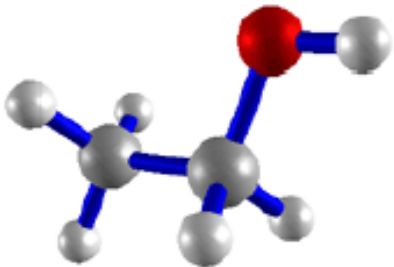
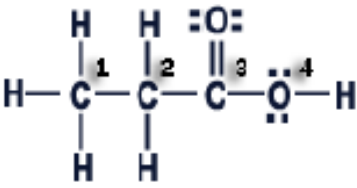
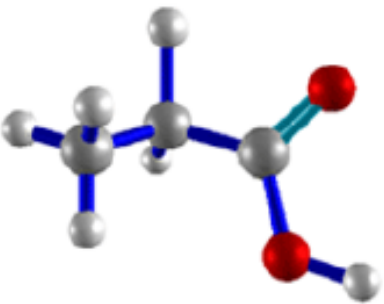
3.10 Molecular Geometries and Bond Angles

Lewis Structure	Classification	X+E	Parent Geometry	Molecular Geometry	Bond Angle	Polarity
<p>CH₄</p>  <p>A: Central atom X: Attachments on A E: Lone pairs on A</p>	<p><u>AX₄E₀</u></p>	<p><u>4</u></p>	<p><i>Electron Pair Geometry</i></p>  <p><u>Tetrahedron</u></p>	 <p><u>Tetrahedron</u></p>	<p><u>~109°</u></p>	
<p>NH₃</p> 	<p><u>AX₃E₁</u></p>	<p><u>4</u></p>	 <p><u>Tetrahedron</u></p>	 <p><u>Trigonal pyramidal</u></p>	<p><u>~109°</u></p>	
<p>H₂O</p> 	<p><u>AX₂E₂</u></p>	<p><u>4</u></p>	 <p><u>Tetrahedron</u></p>	 <p><u>Angular/Bent (~109°)</u></p>	<p><u>~109°</u></p>	

3.10 Molecular Geometries and Bond Angles

Lewis Structure	Classification	X+E	Parent Geometry	Molecular Geometry	Bond Angle	Polarity
H_2CO 	AX_3E_0	<u>3</u>	 <u>Trigonal planar</u>	 <u>Trigonal planar</u>	<u>120°</u>	
NO_2^- 	AX_2E_1 AX_2E_1	<u>3</u> <u>3</u>	 <u>Trigonal planar</u>	 <u>Angular/Bent (120°)</u>	<u>120°</u>	
NO_3^- 	AX_3E_0 AX_3E_0 AX_3E_0	<u>3</u> <u>3</u> <u>3</u>	 <u>Trigonal planar</u>	 <u>Trigonal planar</u>	<u>120°</u>	

3.10 Molecular Geometries and Bond Angles

Lewis Structure	Classification	X+E	Parent Geometry	Molecular Geometry	Bond Angle	Polarity
CO_2 	AX_2E_0	<u>2</u>	 <u>Linear</u>	 <u>Linear</u>	<u>180°</u>	_____
C_2H_4 	1: AX_3E_0 2: AX_3E_0	<u>3</u> <u>3</u>	1: <u>Trigonal planar</u> 2: <u>Trigonal planar</u>		1: <u>120°</u> 2: <u>120°</u>	_____
$\text{C}_2\text{H}_5\text{OH}$ 	1: AX_4E_0 2: AX_4E_0 3: AX_2E_2	<u>4</u> <u>4</u> <u>4</u>	1: <u>Tetrahedron</u> 2: <u>Tetrahedron</u> 3: <u>Tetrahedron</u>		1: <u>$\sim 109^\circ$</u> 2: <u>$\sim 109^\circ$</u> 3: <u>$\sim 109^\circ$</u>	_____
$\text{C}_2\text{H}_6\text{COOH}$ 	1: AX_4E_0 2: AX_4E_0 3: AX_3E_0 4: AX_2E_2	<u>4</u> <u>4</u> <u>3</u> <u>4</u>	1: <u>Tetrahedron</u> 2: <u>Tetrahedron</u> 3: <u>Trigonal planar</u> 4: <u>Tetrahedron</u>		1: <u>$\sim 109^\circ$</u> 2: <u>$\sim 109^\circ$</u> 3: <u>120°</u> 4: <u>$\sim 109^\circ$</u>	_____

3.10 Molecular Geometries and Bond Angles Summary

X+E

ELECTRON PAIR GEOMETRY

MOLECULAR GEOMETRY

4

Tetrahedron ($\sim 109^\circ$)

E=0: Tetrahedron
E=1: Trigonal pyramid
E=2: Angular/Bent $\sim 109^\circ$

3

Trigonal planar (120°)

E=0: Trigonal planar
E=1: Angular/Bent 120°

2

Linear (180°)

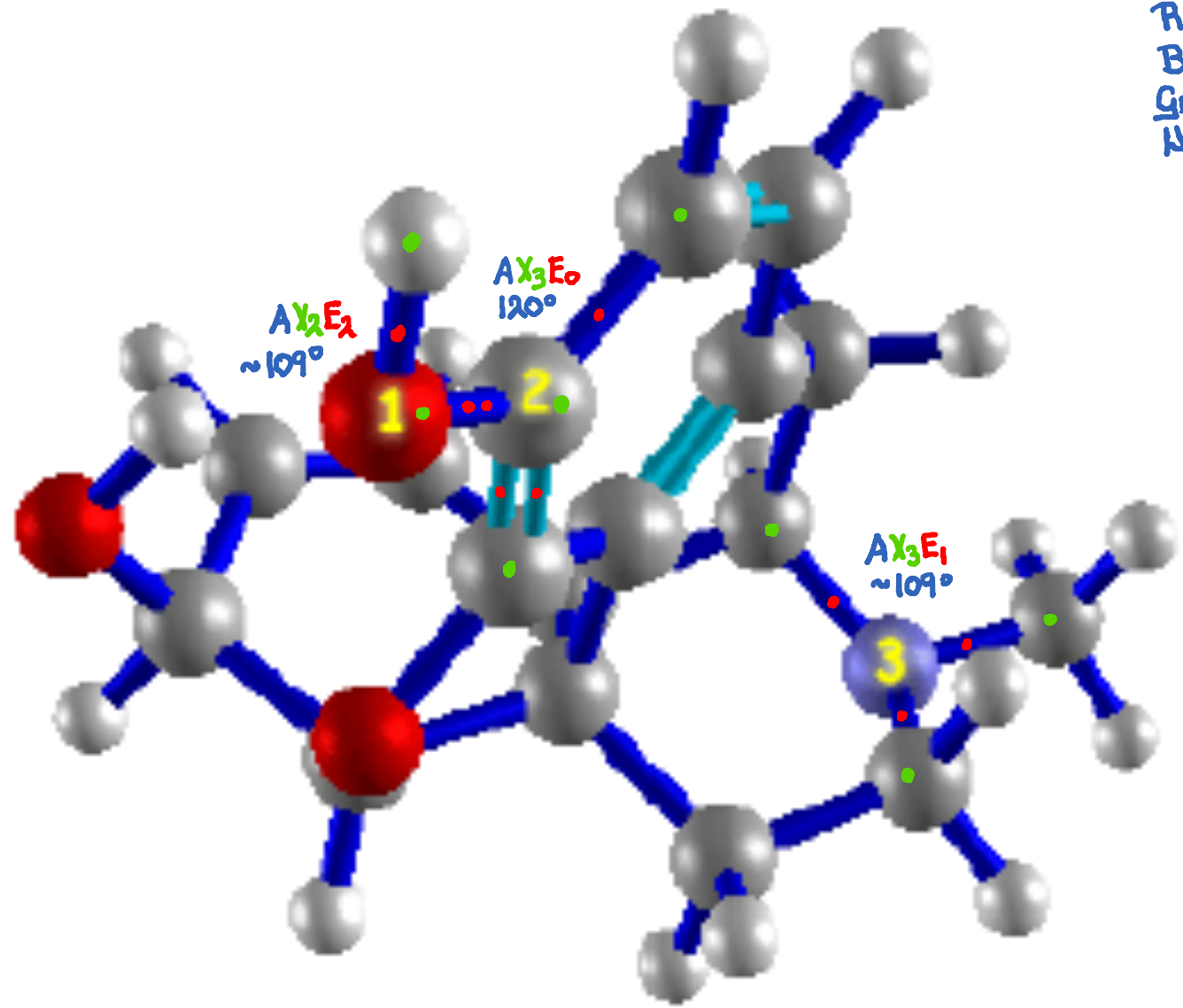
E=0: Linear

3.10 Molecular Geometries and Bond Angles

Morphine

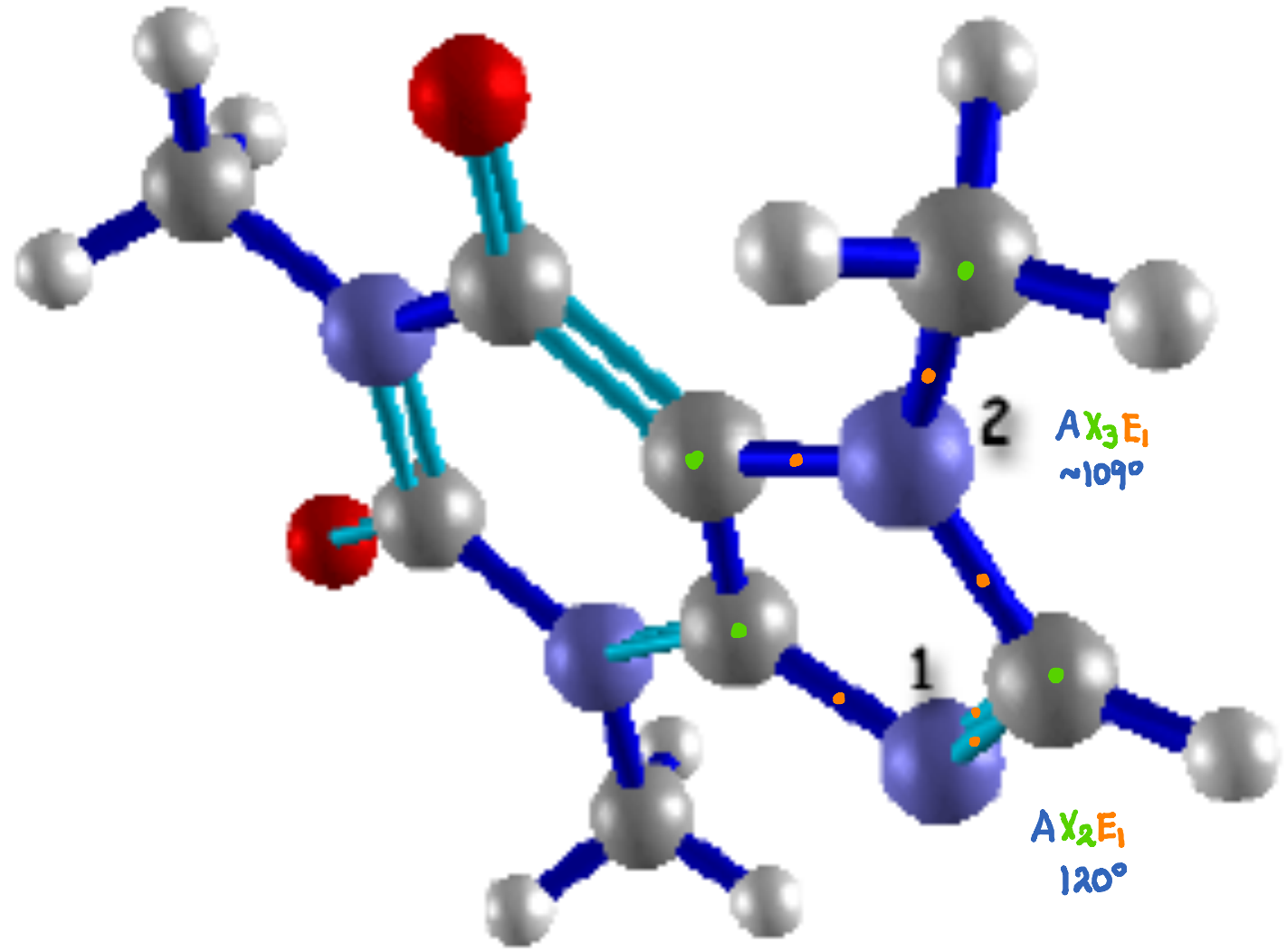
Convention Color Code:

- Red: O
- Blue: N
- Grey: C
- White: H



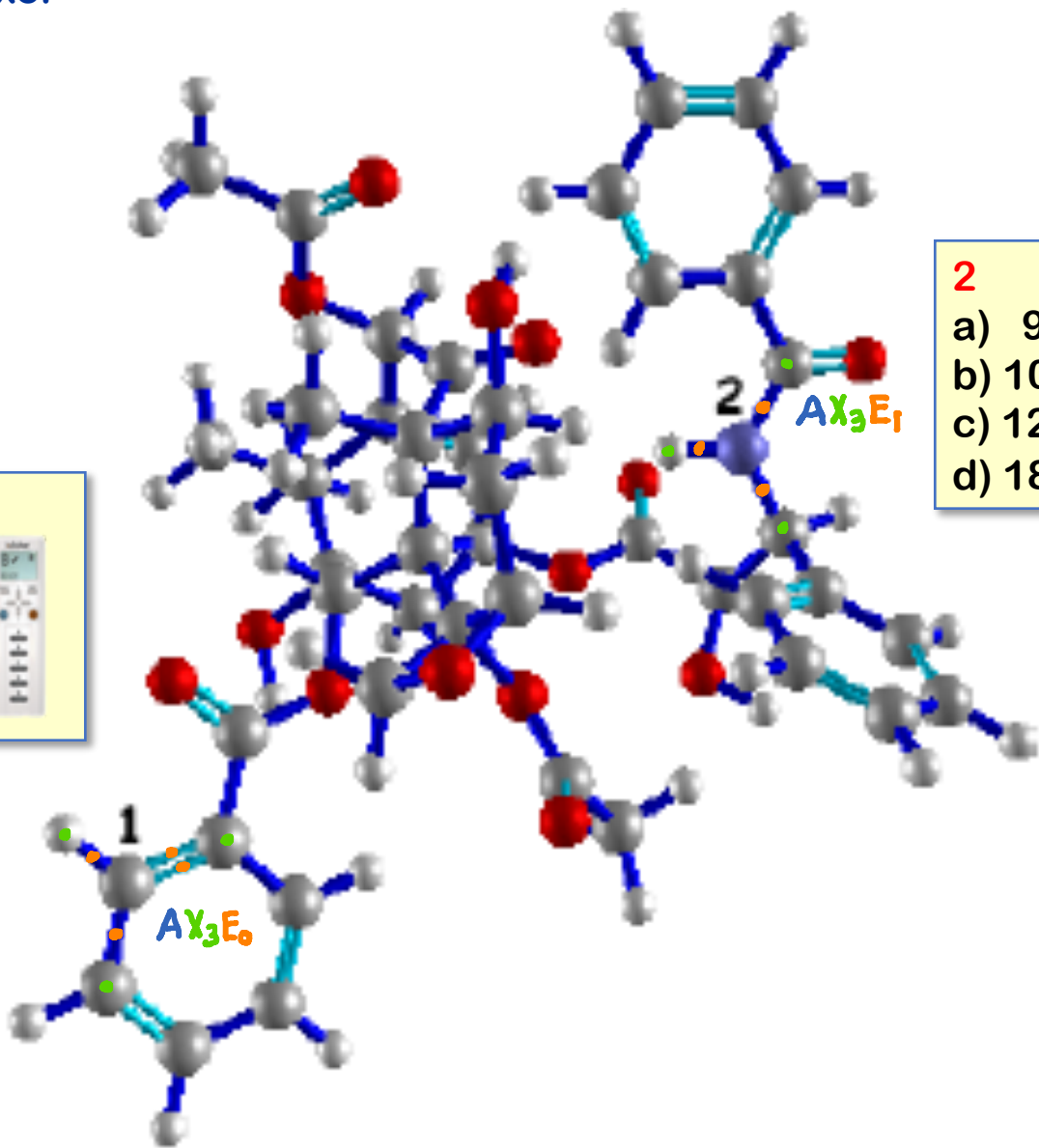
3.10 Molecular Geometries and Bond Angles

Caffeine




3.10 Molecular Geometries and Bond Angles

Taxol




1

- a) 90
- b) 109
- c) 120 ✓
- d) 180



2

- a) 90
- b) 109 ✓
- c) 120
- d) 180




3.10 Molecular Geometries and Bond Angles

Tamiflu

1

- a) 90
- b) 109 ✓
- c) 120
- d) 180



2

- a) 90
- b) 109 ✓
- c) 120
- d) 180

