Announcements – Lecture III – Tuesday, Sep 13th

1. Class Web Site: www.chem.umass.edu/genchem

2. iClicker for credit starts Tuesday , September 20h

   Register your iClicker in Owl (a home work assignment) by Thursday, September 15th

3. First Lab – Saturday, September 24th … 1-4pm … ISB 155 /160 (A-E)

2.4 What Are Atoms Made Of? – *The Three Subatomic Particles*

Name	Symbol	Mass (g)	Charge	Mass*1 (amu)*2
Proton	$^{1}_{1}p$	1.673×10^{-24}	+1	1
Neutron	$^{1}_{0}n$	1.675×10^{-24}	0	1
Electron	$^{0}_{-1}e$	9.109×10^{-28}	-1	0.0005

a) Chemists tend to ignore the mass of the electron.

b) # Protons ... the atom determinator ... #p = Atomic Number (Z).

c) # Neutrons ... the other mass contributor ... #n + #p = Mass Number (A).

d) # Electrons ... determines the charge on the atom.

$^{A}_{Z}X$ X = Symbol

A : Mass Number

Z : Atomic Number

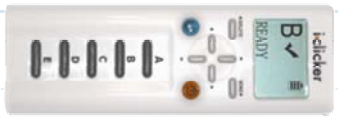
*1: Rounded to 1 sig fig

*2: 1 amu = 1.6605×10^{-24} g

2.4 What Are Atoms Made Of? – *The Three Subatomic Particles*

2.4 Example_1

Which if any of the following species has the same number of Neutrons as it does Electrons?



a) $^{47}_{24}Cr$ ✓ b) $^{24}Mg^{2+}$ c) $^{59}_{27}Co^{2+}$
d) $^{35}Cl^{-}$ e) $^{125}_{50}Sn$

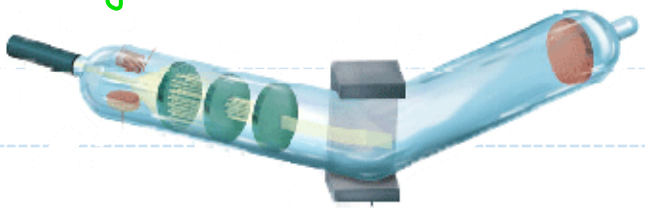
	# Protons	# Neutrons	# Electrons	
$^{47}_{24}Cr$	24	23	24	
$^{24}Mg^{2+}$	12	12	10	
$^{59}_{27}Co^{2+}$	27	32	25	
$^{35}Cl^{-}$	17	18	18	✓
$^{125}_{50}Sn$	50	75	50	

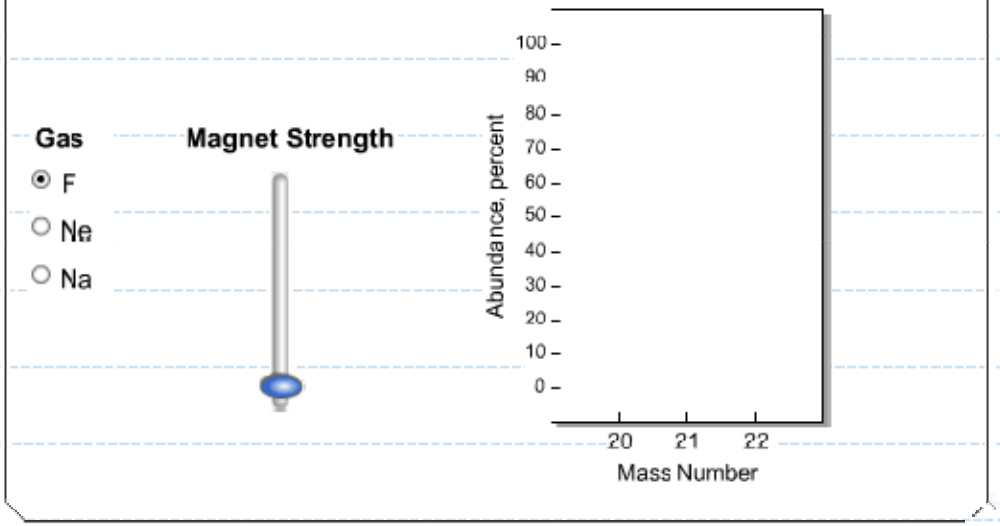
2.4 What Are Atoms Made Of? – *Isotopes*

Isotope : Atoms with the same number of protons but different number of neutrons

Active Figure 2.04: Mass Spectrometer

Interactive figure on class web site





	#p	#n	#e
^{12}C	6	6	6
^{14}C	6	8	6

2.4 What Are Atoms Made Of? — *Atomic Weight*

2.4 Example_2

Chlorine has two naturally occurring isotopes:

^{35}Cl, 75.77% Abundant, Exact Mass 34.96885 amu

^{37}Cl, 24.23% Abundant, Exact Mass 36.96590 amu

What is the Atomic Weight of Chlorine?

Atomic Weight: the weighted average of the naturally occurring isotopes.

$0.7577(34.96885) + 0.2433(36.96590) = 35.45271$ amu

2.4 Example_3

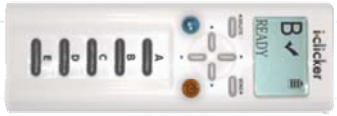
Neon has 3 naturally occurring isotopes:

^{20}Ne, 90.92% Abundant, Exact Mass 19.9989 amu

^{21}Ne, 0.26% Abundant, Exact Mass 20.9975 amu

^{22}Ne, 8.82% Abundant, Exact Mass 21.9979 amu

What is the Atomic Weight of Neon?



The 4th decimal place in the answer is

a) 5 b) 6 c) 7 d) 8✓ e) 9

$0.9092(19.9989) + 0.0026(20.9975) + 0.0882(21.9979) = 20.1778$