

Announcements – Lecture I – Monday, May 20th





University of Massachusetts General Chemistry



	Fall	Spring	Summer
Courses	Chem 102	Chem 101	
	Chem 110	Chem 250	
	Chem 111	Chem 111	Chem 111
	Chem 112	Chem 112	Chem 112
	Chem 121	Chem 122	

2194451

OWL

CRC

Chemistry Dept

Lab Grades

Bus Schedule

Spire

Registrar

Continuing Ed

TA Evaluations

UMail

www.chem.umass.edu/genchem (all lower case)



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Chem 111

Instructors:



Tom Whelan
ISB 241E, 545-6092
whelan@chem.umass.edu

1. [Class Meets](#): MTuWThF 11:15-12:45
2. [Class Location](#): ISB 221
3. [Campus Map](#)



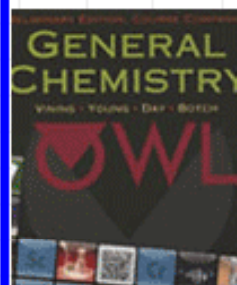
Amanda Hussey
ahussey@chem.umass.edu



Daniel Seeman
dseeman@chem.umass.edu

Summer 2013

Required Materials:



Recommended Textbook:
General Chemistry OWL
Book.

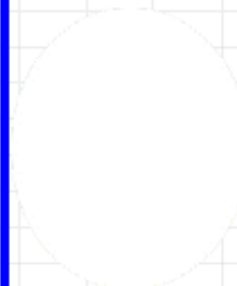
Vining, Young, Day, Botch.

ISBN: 9781285418995

1. [Safety Glasses](#):
2. [Scientific Calculator](#)
(Easy to use and inexpensive)



Max Lein
mlein@chem.umass.edu



Tba
Tba@chem.umass.edu

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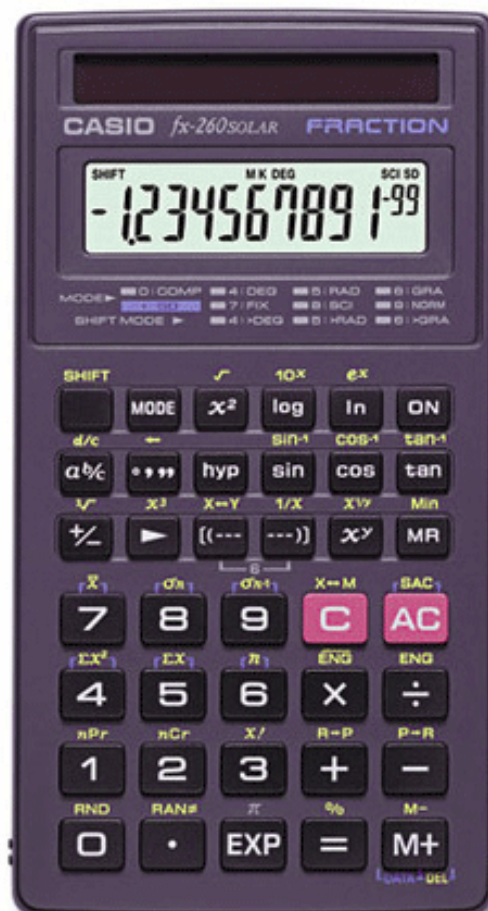
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Syllabus Summary:

→ Short answer format... NON multiple choice... samples on class web site

1. Grading

3 Exams	70%	Highest Exam Score 26%; Lowest 20%; Other 24%
Quizzes	7%	No make-up Quizzes given.
Owl	5%	
Laboratory	18%	

→ 1ST Quiz, Wed, May 22ND
2 LOWEST scores dropped.

Questions taken ~90% of the time
from OWL homework

2. Exam Policies:

You must have a passing exam average in order to pass the course.

3. Lab Policies:

Lab Meets Tuesday/Thursday, 1:30-4:30, ISB 155.

You must complete all of the laboratory experiments to pass the course.

4. Academic Honesty:

You will abide by the academic honesty policy of the campus. I expect you to do your own work on exams and labs. You must flush all calculator memories of any chemistry information before coming to an exam. You MAY NOT bring any additional materials to exams other than a pencil, calculator, and your brain. I take honesty very seriously.

5. Grade Cutoff's:

>90's A. <55 F

The other grade cutoff's will be determined at the end of the semester. However if you want to be assured of a B you should be in the 80's and 70's for a C.

6. Exam Dates:

Exam I	Friday	May	31	ISB 221	11:15-12:45
Exam II	Friday	June	14	ISB 221	11:15-12:45
Exam III	Friday	June	28	ISB 221	11:15-12:45

7. Old Exams

2012	Exam I	Key I	Exam II	Key II	Exam III	Key III
2011	Exam I	Key I	Exam II	Key II	Exam III	Key III
2010	Exam I	Key I	Exam II	Key II	Exam III	Key III
2007	Exam I	Key I	Exam II	Key II	Exam III	Key III

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Date	Daily Schedule ... Updated Wednesday, April 17 :- 05:05 PM																								
Tuesday May 21	<p>Lecture Material:</p> <ul style="list-style-type: none"> 1.4 Unit Conversions <ul style="list-style-type: none"> - How Important Are Units? 2.4 Ionic Compounds <ul style="list-style-type: none"> - Nomenclature ... Some Memorization 2.1 The Structure of the Atom <ul style="list-style-type: none"> - Determining the exact mass of an isotope - Atomic Weight ... What is it? 2.2 Elements and the Periodic Table <ul style="list-style-type: none"> - Divisions in the Periodic Table 2.4 Ionic Compounds → Text book reference <ul style="list-style-type: none"> - Coulomb's Law - Coulomb's Law Data - Coulomb's Law and Solubility <p>Homework:</p> <table border="0"> <tr> <td>OWL</td> <td>Book 1.4</td> <td>Unit Conversions</td> <td>5-22-13</td> </tr> <tr> <td></td> <td>Mastery</td> <td>Unit Conversions</td> <td>5-22-13</td> </tr> <tr> <td></td> <td>Book 2.1</td> <td>The Structure of the Atom</td> <td>5-22-13</td> </tr> <tr> <td></td> <td>Mastery</td> <td>The Structure of the Atom</td> <td>5-22-13</td> </tr> <tr> <td></td> <td>Book 2.2</td> <td>Elements and the Periodic Table</td> <td>5-22-13</td> </tr> <tr> <td></td> <td>Mastery</td> <td>Elements and the Periodic Table</td> <td>5-22-13</td> </tr> </table> <p>Announcements:</p> <ol style="list-style-type: none"> Daily Quizzes -- Start today. 	OWL	Book 1.4	Unit Conversions	5-22-13		Mastery	Unit Conversions	5-22-13		Book 2.1	The Structure of the Atom	5-22-13		Mastery	The Structure of the Atom	5-22-13		Book 2.2	Elements and the Periodic Table	5-22-13		Mastery	Elements and the Periodic Table	5-22-13
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	Mastery	Elements and the Periodic Table	5-22-13																						
Monday May 20	<p>Lecture Material:</p> <ul style="list-style-type: none"> Course Policy and Expectations! <ul style="list-style-type: none"> - Lab Announcements The Computer Resource Center <ul style="list-style-type: none"> - Introduction to Owl A Small Tour of Integrated Sciences Building 1.4 Unit Conversions 																								

Updated twice daily
 ... prior to class
 ... after each class

2.4 Ionic Compounds → Text book reference

Projected coverage ↓





OWL User Login

OWL Login

Login

Login Help



Online Web Learning

University of Massachusetts Amherst Courses - Amherst, Massachusetts
Chemistry General

Login:

NET ID ... See class roster after class if you are UNSURE of this
Use your NetID

Password:

LAST NAME ... or whatever you have used in the past
Use your NetID password

LOG IN

You may safely bookmark this page.



ISB CRC

Integrated Sciences Building Computer Resource Center

[edit]

Chemistry - Biochemistry & Molecular Biology - Biology

ISB Room 325



ISB CRC Collaborative Area

(photo courtesy of Prof. Craig Martin)

The Integrated Sciences Building, which opened in spring 2009, has its own Computer Resource Center dedicated to meeting the needs of students in the College of Natural Sciences and Mathematics who are taking courses in the ISB. As of now that includes the Departments of Chemistry, Biochemistry and Molecular Biology, and Biology.

Information on this page is provisional.

[ISB CRC Software](#)

[Images of the ISB including the CRC](#)

[AV Systems](#) in the CRC (and the rest of the building)

See Also:

[How to find the ISB](#)

SUMMER SCHEDULE

M_F: 8:00am to 4:30pm

1:00 - 2:30pm Manned

by a TA

**see 'Job Schedule' on class
web site for more details.**

[edit]

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} Read prior to first lab ... TUESDAY, May 28th

→ all lab materials posted on class web site

I am retaking this course, do I have to retake the lab?

Here is the scoop on a lab waiver. If you received a grade of **F for the course** then you have to retake the lab. If you received any other grade then you are entitled to a lab waiver. ~~You should print a copy of your grade from Spire and take it to Marie Whalen, in ISB 341 where you will have to fill out some painless paperwork.~~ A cautious note, if your lab **grade is <80**, it is highly recommended that you retake the lab.

→ See me after class.



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Print Page

Note: There are 4 Lab Owls associated with this course. Lab Owl 2-5. The number corresponding to the experiment after which this owl is due.

May 2013				
Monday	Tuesday	Wednesday	Thursday	Friday
20	21	22	23	24
Session I Starts	No Lab	TBA: 1-2:30	No Lab TBA: 1-2:30	TBA: 1-2:30 Add/Drop
27	28	29	30	31
Memorial Day No Class	1:30-4:30 ISB 155 Exp 1	TBA: 1-2:30	1:30-4:30 ISB 155 Exp 2 TBA: 1-2:30	Exam I In Class
June 2013				
Monday	Tuesday	Wednesday	Thursday	Friday
3	4	5	6	7
TBA: 1-2:30	1:30-4:30 ISB 155 Exp 3 Lab Owl 2 Deadline	TBA: 1-2:30	No Lab TBA: 1-2:30	TBA: 1-2:30 Drop with 'W'
10	11	12	13	14
TBA: 1-2:30	1:30-4:30 ISB 155 Exp 4 Lab Owl 3 Deadline	TBA: 1-2:30	No Lab TBA: 1-2:30	Exam II In Class
17	18	19	20	21
TBA: 1-2:30	1:30-4:30 ISB 155 Exp 5 Lab Owl 4 Deadline	TBA: 1-2:30	TBA: 1-2:30	TBA: 1-2:30



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Grading:

Important Summary:

1. **In order to receive a grade in the course you must receive a laboratory grade. Fail the laboratory portion and you fail the entire course, regardless of how you do in lecture. In order to obtain a laboratory grade you must complete ALL the laboratories (see frequently asked question) and made a decent attempt at ALL the assigned laboratory OWLS**
2. **The laboratory grade constitutes $17-20\%$ of the overall course grade.**

Grading within the Laboratory Program:

A final laboratory grade will be posted at the end of the semester before your final exam. This grade is based on the following

Prelab Quiz	25%
Laboratory Reports	45%
Laboratory OWL's	25%
TA Assessment	5%

→ Sample quiz questions.

Disputes/Concerns:

1.4 Unit Conversions

a) Dimensional Analysis

1.4a Example_1

Prior to the metric system, the common unit of weight was the pound (lb). Under the S.I. System, $1 \text{ lb} = 453.5 \text{ g}$. If an old recipe calls for **9 ounces** of flour ($16 \text{ oz} = 1 \text{ lb}$), how many grams of flour is this equivalent to?

$$\frac{9 \text{ ounces}}{16 \text{ ounces}} \times \frac{1 \text{ lb}}{1 \text{ lb}} = 0.5625 \text{ lb}$$

X Cannot do ... NO CONVERSION FACTOR given ... and NO web access!

$$\frac{9 \text{ ounces}}{16 \text{ ounces}} \times \frac{1 \text{ lb}}{1 \text{ lb}} = 0.5625 \text{ lb}$$

$$\frac{0.5625 \text{ lb}}{1 \text{ lb}} \times \frac{453.5 \text{ g}}{1 \text{ lb}} = 255 \text{ g}$$

? What about significant figures?